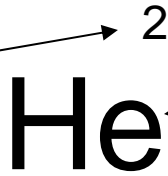


Element Practice

Atomic Number – Represents both the number of protons and electrons in an element.



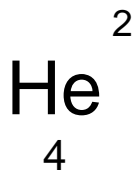
Atomic Symbol – a short hand way of writing each elements name.

Most of the mass comes from the neutron and proton. To find the number of neutrons, subtract the atomic weight from the atomic mass.

4

Atomic Mass – The weight of an element. .

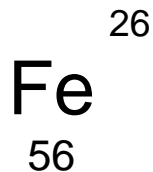
Fill in the necessary information.



element: _____

protons: _____

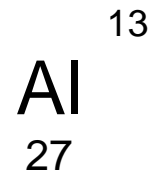
neutrons: _____



element: _____

protons: _____

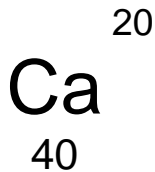
neutrons: _____



element: _____

protons: _____

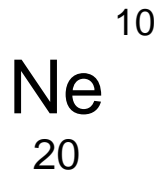
neutrons: _____



element: _____

protons: _____

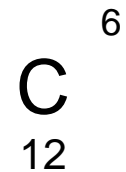
neutrons: _____



element: _____

protons: _____

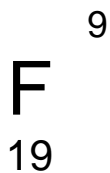
neutrons: _____



element: _____

protons: _____

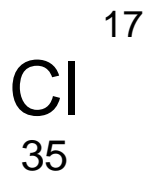
neutrons: _____



element: _____

protons: _____

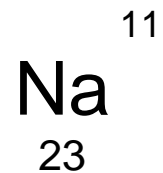
neutrons: _____



element: _____

protons: _____

neutrons: _____



element: _____

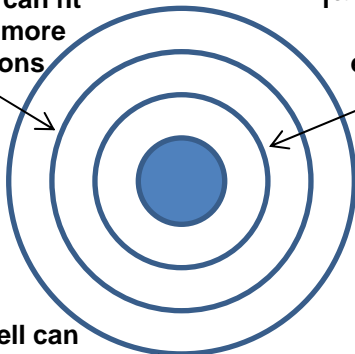
protons: _____

neutrons: _____

Directions: Calculate the number of neutrons and complete the Bohr model for each element.

2nd shell can fit up to 8 more electrons

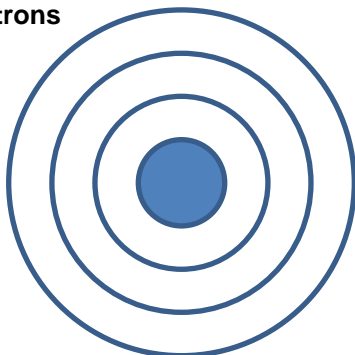
1st shell can fit up to 2 electrons



Element:
Helium

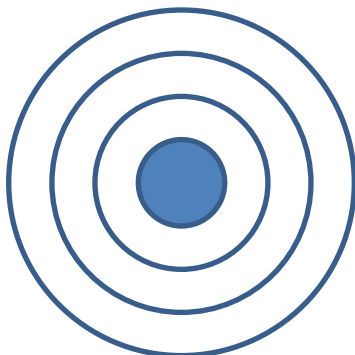
Neutrons:

3rd shell can fit up to 18 more electrons



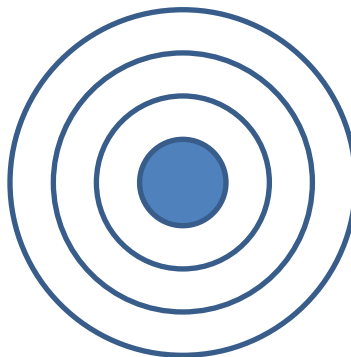
Element:
Beryllium

Neutrons:



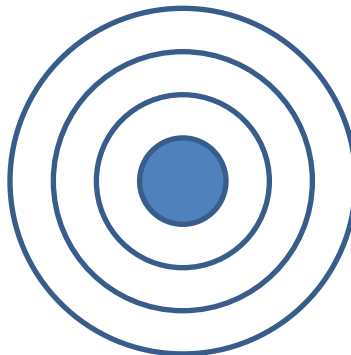
Element:
Boron

Neutrons:



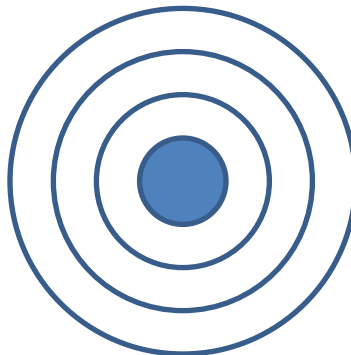
Element:
Neon

Neutrons:



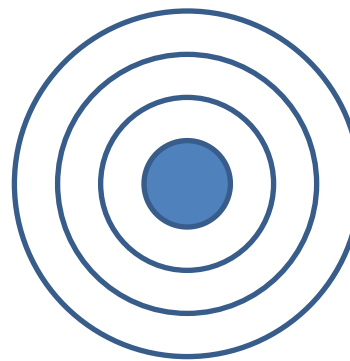
Element:
Oxygen

Neutrons:



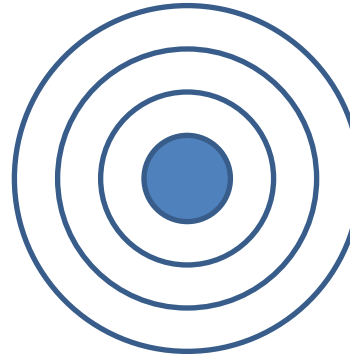
Element:
Sodium

Neutrons:



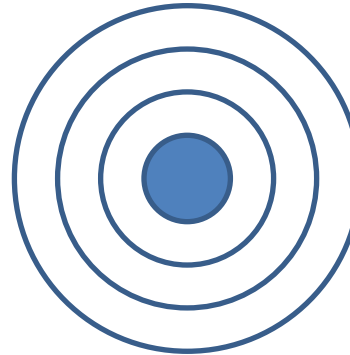
Element:
Nitrogen

Neutrons:



Element:
Chlorine

Neutrons:



Element:
Fluorine

Neutrons:
